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**1. Which statement correct about halogens?**

**(a) They are all diatomic and form univalent ions**

(b) They are all capable of exhibiting several oxidation stables

(c) They are all diatomic and form diatomic ions

(d) They are all reducing agents

**2. Which of the following is not oxidized by $MnO_2$?**

**(a) $F^-$**

(b) $Cl^-$

(c) $Br^-$

(d) $I^-$

**3. Which one of the following is the strongest acid?**

(a) $SO(OH)_2$

(b) $SO_2(OH)_2$

(c) $ClO_2(OH)$

**(d) $ClO_3(OH)$**

**4. Fluorine react with water to give :**

(a) Hydrogen fluoride and oxygen

(b) Hydrogen fluoride and ozone

(c) Hydrogen fluoride and oxygen fluoride

**(d) Hydrogen fluoride, oxygen and ozone**

**5. Astatine is the element below iodine in the group VIIA of the periodic table which of the following statement is not true for astatine?**

(a) It is less electronegative than iodine

**(b) It will exhibit only – l oxidation sate**

(c) Intermolecular forces between the astatine molecule will be large than between iodine molecule

(d) It is compound of diatomic molecule

**6. Which of the following statement is not correct when a mixture of NaCl and $K_2Cr_2O_7$ is gently warmed with con. $H_2SO_4$**

(a) A deep red vapour is evolved.

(b) The vapour when passed in to NaOH solution gives a yellow solution of $Na_2CrO_4$

**(c) Chlorine gas is evolved**

(d) Chromyl chloride is formed

**7. Oxidising action increases from left to right in the order:**

(a) $Cl_2 < Br_2 < I_2 < F_2$

(b) $Cl_2 < I_2 < Br_2 < F_2$

(c) $I_2 < F_2 < Cl_2 < Br_2$

**(d) $I_2 < Br_2 < Cl_2 < F_2$**

**8. The solubility of iodine in water is greatly increased by:**

**(a) $OCl^-$**

(b) $O_2$

(c) $Cl_2$

(d) $Cl^-$

**9. The solubility of iodine in water is greatly increased by :**

(a) Adding an acid

(b) Boiling the solution

(c) Cooling the solution

**(d) Adding potassium iodide**

**10. The correct statement is :**

**(a) $SO_2$ is the anhydride of sulphurous acid**

(b) $H_2S$ is the anhydride of hydrosulphuric acid

(c) $NO_2$ is the anhydride of $HNO_3$

(d) HCL is anhydride of HCl